

Dr Sean's Chemistry Class

Way to SUCCESS

The Periodic Table of the Elements

1 H Hydrogen 1.00794																	2 He Helium 4.003
3 Li Lithium 6.941	4 Be Beryllium 9.012182											5 B Boron 10.811	6 C Carbon 12.0107	7 N Nitrogen 14.00674	8 O Oxygen 15.9994	9 F Fluorine 18.9984032	10 Ne Neon 20.1797
11 Na Sodium 22.989770	12 Mg Magnesium 24.3050											13 Al Aluminum 26.981538	14 Si Silicon 28.0855	15 P Phosphorus 30.973761	16 S Sulfur 32.066	17 Cl Chlorine 35.4527	18 Ar Argon 39.948
19 K Potassium 39.0983	20 Ca Calcium 40.078	21 Sc Scandium 44.955910	22 Ti Titanium 47.867	23 V Vanadium 50.9415	24 Cr Chromium 51.9961	25 Mn Manganese 54.938049	26 Fe Iron 55.845	27 Co Cobalt 58.933200	28 Ni Nickel 58.6934	29 Cu Copper 63.546	30 Zn Zinc 65.39	31 Ga Gallium 69.723	32 Ge Germanium 72.61	33 As Arsenic 74.92160	34 Se Selenium 78.96	35 Br Bromine 79.904	36 Kr Krypton 83.80
37 Rb Rubidium 85.4678	38 Sr Strontium 87.62	39 Y Yttrium 88.90585	40 Zr Zirconium 91.224	41 Nb Niobium 92.90638	42 Mo Molybdenum 95.94	43 Tc Technetium (98)	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.90550	46 Pd Palladium 106.42	47 Ag Silver 107.8682	48 Cd Cadmium 112.411	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.60	53 I Iodine 126.90447	54 Xe Xenon 131.29
55 Cs Cesium 132.90545	56 Ba Barium 137.327	57 La Lanthanum 138.9055	72 Hf Hafnium 178.49	73 Ta Tantalum 180.9479	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.078	79 Au Gold 196.96655	80 Hg Mercury 200.59	81 Tl Thallium 204.3833	82 Pb Lead 207.2	83 Bi Bismuth 208.98038	84 Po Polonium (209)	85 At Astatine (210)	86 Rn Radon (222)
87 Fr Francium (223)	88 Ra Radium (226)	89 Ac Actinium (227)	104 Rf Rutherfordium (261)	105 Db Dubnium (262)	106 Sg Seaborgium (263)	107 Bh Bohrium (262)	108 Hs Hassium (265)	109 Mt Meitnerium (266)	110 (269)	111 (272)	112 (277)	113	114				
		58 Ce Cerium 140.116	59 Pr Praseodymium 140.90765	60 Nd Neodymium 144.24	61 Pm Promethium (145)	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.92534	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93032	68 Er Erbium 167.26	69 Tm Thulium 168.93421	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.967		
		90 Th Thorium 232.0381	91 Pa Protactinium 231.03588	92 U Uranium 238.0289	93 Np Neptunium (237)	94 Pu Plutonium (244)	95 Am Americium (243)	96 Cm Curium (247)	97 Bk Berkelium (247)	98 Cf Californium (251)	99 Es Einsteinium (252)	100 Fm Fermium (257)	101 Md Mendelevium (258)	102 No Nobelium (259)	103 Lr Lawrencium (262)		

Dr Sean Chemistry

DSE Challenge by
Dr Sean Chemistry

Topic 3 Mole Calculation



You have 20 mins to complete 10 MC Questions.

1. Calculate the mass of 0.25 mol of $\text{FeCl}_2 \cdot 5\text{H}_2\text{O}$.
(Relative atomic mass, Fe: 55.8, Cl: 35.5, O: 16.0, H: 1.0)
Level 3

A. 31.7 g
B. 54.2 g
C. 507 g
D. 867 g

2. Calculate the number of formula units in 0.301 g of MgSO_4 .
(Relative atomic mass, S: 32.1, Mg: 24.3, O: 16.0,
Avogadro's constant, $6.02 \times 10^{23} \text{ mol}^{-1}$)
Level 3

A. 1.505×10^{21}
B. 2.503×10^{21}
C. 1.448×10^{26}
D. 2.408×10^{26}

3. What is the empirical formula of a compound with the following percentage by mass:

Na: 33.3 % N: 20.3% O: 46.4%
(Relative atomic mass, Na: 23.0, O: 16.0, N: 14.0)
Level 3

A. Na_2NO_2
B. Na_2NO_3
C. NaNO_2
D. NaNO_3

4. Calculate the maximum mass of lithium oxide that can be obtained when 1.725 g of lithium is heated with oxygen.
(Relative atomic mass, O: 16.0, Li: 6.9)
Level 3

A. 2.86 g
B. 3.73 g
C. 7.45 g
D. 14.9 g

5. Calculate the number of Fe^{3+} in 1.20 g $\text{Fe}_2(\text{SO}_4)_3$.
(Relative atomic mass, S: 32.1, Fe: 55.8, O: 16.0, Avogadro's constant, $6.02 \times 10^{23} \text{ mol}^{-1}$)
Level 4-5

A. 1.81×10^{21}
B. 3.48×10^{21}
C. 3.61×10^{21}
D. 6.96×10^{21}

6. 21.0 g of an organic compound, composed C, H and O only was burnt completely to give 30.8 g of carbon dioxide and 12.6 g of water, determine the empirical formula of the organic compound.
(Relative atomic mass, C: 12.0, H: 1.0, O: 16.0)
Level 5

A. CHO
B. CH_2O
C. C_2HO_2
D. Cannot be determined

7. Calculate the maximum mass of magnesium oxide that can be formed by heating 9.72 g of magnesium with 9.60 g of oxygen.

(Relative atomic mass, Mg: 24.3, O: 16.0)

Level 5

A. 8.06 g

B. 12.1 g

C. 16.1 g

D. 24.2 g

8. Calculate the maximum mass of aluminum oxide that can be obtained by heating 13.5 g of aluminum with 9.60 g of oxygen.

(Relative atomic mass, Al: 27.0, O: 16.0)

Level 5

A. 20.4 g

B. 25.5 g

C. 30.6 g

D. 51.0 g

9. 14 g of N_2 contains X atoms, how many atoms does 8 g O_2 contain in terms of X?

(Relative atomic mass, N: 14.0, O: 16.0)

Level 5*

A. $X/2$

B. X

C. 2X

D. 4X

10. A magnesium chloride sample contains 1.49×10^{22} Cl^- ion and occupies a volume of 0.678 cm^3 . Calculate the density of the magnesium chloride sample.

(Relative atomic mass, Cl: 35.5, Mg: 24.3, Avogadro's constant, $6.02 \times 10^{23} \text{ mol}^{-1}$)

Level 5**

- A. 0.575 g cm^{-3}
- B. 1.74 g cm^{-3}
- C. 3.48 g cm^{-3}
- D. Cannot be determined

🔥 想知道你嘅程度去到邊？ 🔥

✅ 用Google Form 交答案

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